

Chemistry: Acid-Base Equilibrium

Chapter 1: Acid-Base Equilibrium

1.1 Acid-Base Equilibrium

Acid-base equilibrium is a chemical reaction between an acid and a base. The acid donates a proton (H^+) to the base, forming a conjugate base and a conjugate acid. The equilibrium constant, K_a , is a measure of the strength of the acid. The larger the K_a , the stronger the acid.

1.2 Acid-Base Equilibrium

The acid-base equilibrium is a chemical reaction between an acid and a base. The acid donates a proton (H^+) to the base, forming a conjugate base and a conjugate acid. The equilibrium constant, K_a , is a measure of the strength of the acid. The larger the K_a , the stronger the acid.

1.3 Acid-Base Equilibrium

The acid-base equilibrium is a chemical reaction between an acid and a base. The acid donates a proton (H^+) to the base, forming a conjugate base and a conjugate acid. The equilibrium constant, K_a , is a measure of the strength of the acid. The larger the K_a , the stronger the acid.

1.4 Acid-Base Equilibrium

The acid-base equilibrium is a chemical reaction between an acid and a base. The acid donates a proton (H^+) to the base, forming a conjugate base and a conjugate acid. The equilibrium constant, K_a , is a measure of the strength of the acid. The larger the K_a , the stronger the acid.

1.5 Acid-Base Equilibrium

The acid-base equilibrium is a chemical reaction between an acid and a base. The acid donates a proton (H^+) to the base, forming a conjugate base and a conjugate acid. The equilibrium constant, K_a , is a measure of the strength of the acid. The larger the K_a , the stronger the acid.